

Department of Chemistry  
Kisan Veer Mahavidyalaya, Wai  
Date: 02.09.2018

## NOTICE

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All the Students of B.Sc. III are hereby informed that your Seminar on Inorganic Chemistry by Aishwarya B. Gadhave on Pearsons's HSAB Concept will be conducted on Tuesday ,03/09/2018 at 9:00 am So all should remain present for the same.



  
Head  
Department Of Chemistry  
Kisan Veer Mahavidyalaya, Wai

Department Of Chemistry  
Kisan Veer Mahavidyalaya, Wai

Seminar Attendance sheet

B.Sc.III 03/09/2018

Sr.No.	Roll No.	Name Of Students	Signature	Teacher Name
1	07	Ombale Jyoti R.	J.R. Ombale	Dr. P.H. Bhoite
2	8	Snehal S. Pisal	Snehal	
3	04	Mayuri Dudhe R	M. Dudhe	
4	10	Bhosale Pooja M.	P. Bhosale	
5	12	Shweta Bhosale S.	S. Bhosale	
6	17	Poonam Dhebe	P. Dhebe	
7	20	Karpe Sayali	K. Karpe	
8	31	chavan Rajesh	R. Chavan	
9	26	Shubhangi Shinde	S. Shinde	
10	34	Dhawan Rishikesh.	R. Dhawan	
11	37	Shelar Deepali	D. Shelar	
12	41	Gholap Pratik	P. Gholap	
13	44	Akshay S. Malusare	A. Malusare	
14	51	Pisal Pooja U.	P. Pisal	
15	57	Kasare Madhumala.	M. Kasare	
16	23	Gadhare Aishwarya B.	A. Gadhare	
17	148	Wagh Smit	S. Wagh	
18	63	Yadav Prajakta	P. Yadav	
19	177	Dhumal Nikita	N. Dhumal	
20	222	Cheran Rukan	R. Cheran	
21	165	Babar Ambika	A. Babar	
22	181	Snehal Kadam	S. Kadam	
23	192	Jagtap Sayali S.	S. Jagtap	
24	197	Jagtap Pragati	P. Jagtap	
25	200	Kadam Shubhem	S. Kadam	
26	205	Mahangade Akash.	A. Mahangade	
27				
28				
29				
30				

Male = 08

Female = 18

Total = 26



  
Head

Department Of Chemistry  
Kisan Veer Mahavidyalaya, Wai

Janata Shikshan Sanstha's  
Kisan Veer Mahavidyalaya, Wai  
Department Of Chemistry  
Seminar B.Sc. III  
Report

A seminar of Aishwarya Gadhave a student of B.Sc.III was conducted on Tuesday, 03/09/2018 at 9:00 am on Pearson's HSAB Concept. The points included in her seminar are as follows-

In 1963, R.G. Pearson extended & generalized the qualitative correlation between Lewis acids & Lewis bases by classifying them into two categories: Hard & soft.

i) class a - The class 'a' metals which are small & less polarisable, prefer to combine with non-metals or ligands which are also small & not very polarisable. Pearson called such metals as hard acids & the corresponding ligands as soft bases.

ii) class b - have large size, more or easily polarisable, prefer to combine with non-metals or ligands having similar properties. Pearson called such metals as soft acids & the ligands as soft bases.

Principle -

"Hard acids prefer (tend) to combine (co-ordinate) with hard bases & soft acids prefer to bind with soft bases & give stable complex compounds."

It should be noted that the statement given above is not a theory or explanation but it is simple rule of thumb which enables us to predict the relative stabilities of acid-bases adducts qualitatively.

Department of Chemistry  
Kisan Veer Mahavidyalaya, Wai  
Date: 04.09.2018

## NOTICE

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All the Students of B.Sc. III are hereby informed that your Seminar on Inorganic Chemistry by Bhosale Shweta on Merits and Demerits of MOT will be conducted on Thursday ,05/09/2018 at 9:00 am So all should remain present for the same.



  
Head  
Department Of Chemistry  
Kisan Veer Mahavidyalaya, Wai

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Janata Shikshan Sanstha's  
Kisan Veer Mahavidyalaya, Wai  
Department Of Chemistry  
Seminar B.Sc. III  
Report

A seminar of Shweta Bhosale a student of B.Sc.III was conducted on Tuesday, 05/09/2018 at 9:00 am on Merits and Demerits of MOT. The points included in her seminar are as follows-

I] Merits of MOT —

- 1) MOT accounts for all possible interactions between metal orbitals & the ligand orbitals, including the electrostatic situation, are thoughtfully considered.
- 2) It successfully accounts the presence of  $\pi$  bonding & the increased stability of complexes especially with strong field ligands.
- 3) It gives information about high spin & low spin complexes, their different magnetic moments & variation in  $\Delta_0$  values.
- 4) MOT account for stability, geometry & relative energies of different structures of metal complexes, unlike CBT or CFT.

II] Demerits of MOT —

- 1) Since calculations & predictions are made only with great difficulty, it is not that easy to apply rigorously to different problems.
- 2) Many approximations are to be made in its quantitative application to multi-electron multi-atom complex ion systems.
- 3) MOT does not lend itself to pictorial representation of bonding.
- 4) Many times its qualitative application to the simple regular Oh. complexes without  $\pi$  bonding, is only taken into consideration, on the basis of LCAO approximation. In short, MOT is too general to apply.

Department Of Chemistry  
Kisan Veer Mahavidyalaya, Wai

Seminar Attendance sheet

B.Sc.III 05/09/2018

Sr.No.	Roll No.	Name Of Students	Signature	Teacher Name
1	04	Dudhe Mayuri R.	(Dudhe)	Dr. P.H. Bhoirle
2	10	Bhasale Pooja M.	(Bhasale)	
3	12	Shweta Bhasale S.	(Shweta)	
4	15	Pranjana Barkade.	(Pranjana)	
5	07	Ombale Jyoti R.	(J.R. Ombale)	
6	8	Snehal S. Pisal.	(Snehal)	
7	20	Karpe Sayali	(Karpe)	
8	41	Gholap Pankaj	(Gholap)	
9	31	Chavan Rajesh	(Chavan)	
10	26	Shubhangi Shinde	(Shubhangi)	
11	17	Poonam Dhebe	(Poonam)	
12	37	Shelar Deepali	(Shelar)	
13	34	Dhawan Rishikesh	(Dhawan)	
14	44	Akshay S. Malysare	(Akshay)	
15	51	Pisal Pooja U.	(Pisal)	
16	63	Yadav Prajakta	(Yadav)	
17	57	Vasave Madhumata	(Vasave)	
18	23	Godhave Akshwary B	(Godhave)	
19	146	Wagh Sujit.	(Wagh)	
20	165	Babar Ankita.	(Babar)	
21	177	Dhumal Nikita	(Dhumal)	
22	181	Snehal Kadam	(Snehal)	
23	222	Charan Rokan	(Charan)	
24	192	Jagtap Sayali S.	(S.S. Jagtap)	
25	197	Wagh Pragati	(Wagh)	
26	200	Yadav Shubham	(Yadav)	
27	205	Mahangade Akash	(Mahangade)	
28				
29				
30				

Male = 08

Female = 19

Total = 27



*(Signature)*  
Head

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